Good, but could be better

	Stages	Teacher-student exchanges
Teacher	Prepare sentence	Please pay attention to the first two sentences under the section B as well as the chemical equation below. These two sentences describe what reactants and products are in a neutralisation reaction 'the acid and alkali have undergone a reaction. They are called reactants. A salt and water are formed in the reaction. They are called the products.'
Student Teacher	Focus Identify Affirm	[student name] What are the reactants in a neutralization reaction? Acid and alkali. Exactly.
Student Teacher	Focus Identify Affirm	[student name] What are the products in a neutralization reaction? Salt and water. Right.
Teacher		Once again have a look at the chemical equation in section B. Neutralization occurs when acid and an alkali react with each other to form a salt and water. [point at the equation] The reactants ate shown on the left side, and the products are shown on the right side.
	Elaborate	For example, when we brush our teeth with toothpaste, the alkali in the toothpaste reacts with the acids produced by the bacteria in our mouth. They are converted into a salt and water, so our teeth are cleaned. This is one application of neutralisation mentioned in the "key concept highlight" box.

Improved from the first example

	Stages	Teacher-student exchanges
Teacher	Prepare	Look at the first two sentences under
		the section B. OK? Section B, first two
		sentences. (switches to Chinese; with
		some English words underlined)
		These two sentences describe what
		reactants and products are in a
		neutralization reaction. "Acid and alkali
		have undergone a reaction. They are
		<u>called reactants. A salt and water are</u>
		formed in the reaction. They are called
		the products."
		Goes to board and writes:
		"REACTANTS: acid + alkali. Arrow
		PRODUCTS: salt and water."
		(switches to Chinese) Understand?
		When acid and alkali react, they form
		salt and water.
		(writes in English on board the
		nominalization)
		- "The <u>reaction</u> (N) of acid and alkali
		produces (VB) salt and water."
		- "When acid and alkali <u>react</u> (VB), the
		products (N) are salt and water."
		(explains in English) OK? Reaction /
		react. Product / produce. Please pay
		attention to the nouns and verbs.
		(switches to Chinese) OK, I want to
	Focus	check you understood. (switches to
		English) What are the reactants in a
	Identify	neutralization reaction?
Student		Acid and alkali
Teacher	Affirm	Exactly. (switches to Chinese, with one
		English word) And when they react,
		they <u>neutralize</u> each other. OK? They
		take the "bite" out of one another.

	Focus	(switches to English) They neutralize one another. We call this "neutralization reaction." (writes on board above what she previously wrote: "NEUTRALIZATION REACTION.") Let me check your understanding. What are the products from this
	Identify	neutralization reaction?
Student		Salt and water.
Teacher	Affirm, then Focus	Right. (points to salt and water on the board)
Teacher	Repeat/Reiterate	Once again, please look at the chemical equation in section B. (points to the handout in her hands; starts to read) "Neutralization occurs when acid and an alkali react with each other to form a salt and water." (switches to Chinese) Please point to where I am in the reading. (walks around to check students are pointing correctly) Good. See, the reactants are shown on the left side, and the products are shown on the right side. (switches to English) I repeat: reactants are shown on the left, products on the right.
	Elaborate	Now, can you think of a reaction like this in daily life? For example, when we brush our teeth with toothpaste, the alkali in the toothpaste reacts with the acids in our mouth. We have bacteria, xìjùn, in our mouth which produce acid. Yes, bacteria, xìjùn, produce acid. And acid reacts with the alkali in the toothpaste. In the neutralization reaction the alkali and acid are converted into salt and water, so our teeth are cleaned. (pauses) This is one application of neutralization mentioned in the "key concept highlight" box. See if you can find that box on the page.